CC14 Rates of Reaction

CC14a Rates of reaction

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe different changes that can occur as a reaction proceeds.			
70)	Suggest different experimental methods to investigate rates of reaction (e.g. measurements of mass of reactants against time, volume of gas released against time, concentration of reactant or product against time).			
7th	Use graphs of changes (in mass, volume or concentration of reactant or product) against time, to interpret what is happening during reactions.			

CC14b Factors affecting reaction rates

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Explain what has to happen for reactions to take place.			
9th	Explain why changes in the energy of particles affect rates of reaction.			
91	Explain why changes in the frequency of collisions between particles affect the rate of reaction.			
910	Explain why changes in temperature, concentration, surface area and pressure affect the rate of reaction (surface area for solids, pressure for gases only).			
8 th	Describe ways of speeding up or slowing down chemical reactions.			

CC14c Catalysts and activation energy

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe what a catalyst does.			
7 th	Explain how catalysts are useful.			
8 th	Explain what the activation energy of a reaction is.			
9th	Explain how catalysts speed up chemical reactions.			
7 th	Describe what enzymes are.			
6 th	Name one or more examples of enzymes.			